

THERMODYNAMIC MODELING OF COGENERATION MINI CHP USING AIR CONVERSION OF DIESEL FUEL AND ELECTROCHEMICAL GENERATOR

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ABSTRACT

At present, the production of electrical and heat power uses diesel-generator technology with a limited service life of engines and extremely low efficiency of the expensive fuel used. In this paper, an innovative technology has been considered for the combined electrical and heat power production using the preliminary conversion of diesel fuel into synthesis gas with its subsequent supply to a high temperature electrochemical generator (ECG). Synthesis gas for the operation of the electrochemical generator was produced by air conversion of motor diesel fuels in a catalytic burner reactor. On the basis of heat balances of the burner, ECG and waste-heat boiler-utilizer, electrical efficiency of the solid oxide fuel cells' (SOFC) battery, chemical efficiency of the burner, the temperature at the SOFC anode, the EMF of the planar cell, a portion of hydrogen oxidized at the SOFC anode, specific consumption of diesel fuel for the production of electrical and heat power were calculated. Specific consumption of diesel fuel for the production of electrical and heat power was found to be equal to 114 g/kWh (162 g r.f./kW-h) and 31.7 kg/GJ (45.1 kg r.f./GJ, 189 kg r.f./ Gcal), respectively. Specific fuel consumption is similar to an up-to-date CHP and is significantly lower than the consumption of modern diesel-electric stations of equal power.

Keywords: diesel fuel, electrochemical generator, fuel cell, stoichiometry.

1 INTRODUCTION

The achievements made over the past several decades in the development of efficient high-temperature solid oxide fuel cells (SOFCs) capable of generating electrical energy from hydrogen contained in the synthesis gas have led to a rapid progress in studies on conversion of natural (methane, coal, petroleum derivatives, etc.) and synthetic (methanol, ethanol, etc.) fuels into synthesis gas followed by its use in SOFCs [1–15]. In [16], more than 220 variants of the SOFC operation with recirculation of anode gases and without it, using methane as a fuel and air as an oxidant were calculated. It was shown that the fuel utilization factor of high-temperature SOFCs can reach from 60% to 80%, while the choice of the scheme has little effect on the efficiency. In [17], an experimental analysis of a pilot electrochemical generator (ECG) installation based on SOFC with a power of 5 kW was presented. Natural gas was used as a fuel; air was used as an oxidant. The fuel utilization ratio reached 70%. The installation has worked for more than 7000 hours. In [18], the power plant based on the SOFC battery with recirculation of anode gases was investigated, where methane was used as a fuel and water as an oxidant. The fuel utilization factor reached 60%.

In [19], the performance characteristics of a 5 kW power plant were investigated. Natural gas was used as a fuel; water was used as an oxidant. The efficiency of the SOFC battery was 73%, the electric efficiency of ECG was 51%. The fuel utilization factor was 52.6%. Calculation studies of the power plant with an electric power of 7.3 kW, operating on coal of the

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Borodinsky coal strip mine with water as an oxidant, were presented in [20]. The fuel utilization ratio was 36.4%, the efficiency of the SOFC battery was 36.4%, the electric efficiency of the ECG (gross) was 13.2%, and the chemical efficiency of the gas generator was 36.4%. Specific consumption of reference fuel for the production of electric power was equal to 0.338 kg r.f./kWh, and heat energy supplied to heat networks was 93 kg r.f./GJ. Calculation studies were carried out for the ECG, where natural gas was used as a fuel and water was used as an oxidant [21]. The electric power (Q_e) of the power plant was 5.33 kW. The specific fuel consumption for the generation of electrical power was 0.16 kg r.f./kWh, and for the heat power, 44.7 kg r.f./GJ at a fuel utilization ratio equal to 0.71.

Ref. [22] was devoted to theoretical studies of power parameters of the power plant on the basis of SOFC using solid municipal waste as a fuel and water vapor as an oxidant. The specific consumption of reference fuel for the production of electrical power was 0.153 kg r.f./kW-h, and heat power, 42.8 kg r.f./GJ. In order to increase the ECG efficiency, it is recommended in some cases to create hybrid installations using a gas turbine [23]. The electrical efficiency of the power plant was 66% and that of the SOFC battery was 61.5%.

In this paper, an innovative technology has been considered for the combined electrical and heat power production using the preliminary conversion of diesel fuel into synthesis gas with its subsequent supply to a high temperature electrochemical generator (ECG)

2 MODELLING OF PHYSICOCHEMICAL PROCESSES IN THE UNITS OF THE MINI-CHPP BASED ON THE CATALYTIC BURNER FOR AIR CONVERSION OF MOTOR DIESEL FUEL AND ELECTROCHEMICAL GENERATOR

2.1 Principle and technological model of the process

In [24] more than 220 variants of the SOFC operation with recirculation of anode gases and without it have been calculated. It has been concluded that the choice of the scheme has little effect on the process efficiency, thus it is advisable to choose the simplest scheme. In the scheme proposed, the anode gases recirculation is not considered. Planar fuel cells with direct flow of fuel and oxidizer are used. Catalysts at the anode and cathode are made of Ni-YSZ and $(\text{ZrO}_2)_{0.9}(\text{Sc}_2\text{O}_3)_{0.1}$ is used as an electrolyte (see Fig. 1).

Motor diesel fuel is supplied to the waste-heat boiler (WHB) for heating to boiling and evaporation and enters the catalytic burner in a vaporized form. There also comes the air heated in WHB. When the air flow rate (a dimensionless ratio of actual amount of air fed into the combustion chamber to the theoretically required for burning) $\alpha = 0.4$, diesel conversion occurs with the formation of synthesis gas. Then the synthesis gas is cooled from 1284°C to 800°C in the cooler by the air supplied to the cathode channel.

Synthesis gas is supplied to the anode channel. At the ECG anode, a share of the hydrogen is oxidized by the oxygen of the air supplied to the cathode channel. Products from the anode channel and vitiated air from the cathode channel are fed to the waste-heat boiler, where the remaining hydrogen and carbon monoxide are oxidized. The heat of oxidation is used for heating the primary air and evaporation of the diesel fuel. Then the combustion products from WHB enter the network heater to heat the network water for the needs of hot water supply to heat consumers. Combustion products after the network heater are removed through the stove chimney at a temperature of 120°C.

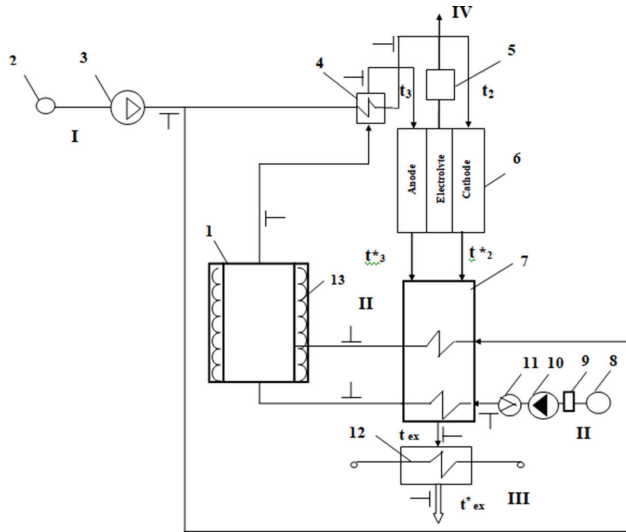


Figure 1: Schematic diagram of mini-CHPP on products of air conversion of motor diesel fuel and electrochemical generator:

- 1 – catalytic burner; 2 – air pipe; 3 – compressed-air blower; 4 – synthesis gas cooler; 5 – electrical customer; 6 – electrochemical generator based on SOFC; 7 – waste-heat boiler-utilizer; 8 – motor diesel fuel manifold; 9 – fuel filter; 10 – fuel pump; 11 – flow-sensing element; 12 – network heater; 13 – thermal insulation.
- I – air input; II – motor diesel fuel input; III – output of combustion products; IV – electrical power output.

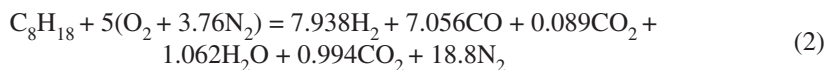
2.2 Modelling of the processes in the technological scheme units

2.2.1 Modelling of the processes in the catalytic burner

It is known [25] that motor diesel fuel contains (in weight %) carbon – 87, hydrogen – 12.6 and oxygen – 0.5. For further calculations let us define the diesel fuel composition in atomic fractions of the elements contained in it [25]. It contains $n = 87/12 = 8$ of carbon, $m = 2n + 2 = 18$ of hydrogen and $l = 0.5/16 = 0.03$ of oxygen. Since the oxygen content is low, it can be neglected in further calculations. In this case, the diesel fuel formula can be written as C_8H_{18} . The minimum coefficient of air consumption at which the emission of soot in the products of conversion is thermodynamically impossible:

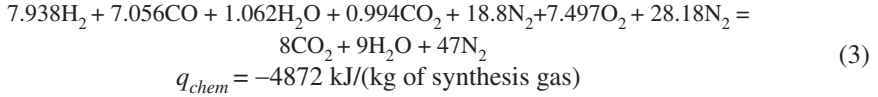
$$\alpha = (n - 1) / (2n + 0.5m - 1) \tag{1}$$

We take the coefficient of air flow $\alpha > \alpha_c$ ($\alpha = 0.4$). When, in the reaction volume of the catalytic burner, the air flow rate is equal to 0.4, synthesis gas is produced according to the stoichiometric equation:

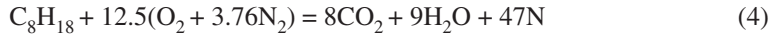


The composition of the synthesis gas produced (in volume %): $H_2 - 22.1$; $CO - 19.68$; $H_2O - 2.96$; $CO_2 - 2.77$; $N_2 - 52.49$.

The heat of combustion of the synthesis gas is determined according to the stoichiometric equation of complete oxidation of synthesis gas:



Complete combustion of the diesel fuel at the air flow rate factor α equal to 1 is described by the stoichiometric equation:



The number of kilomoles of air at the flow rate factor $\alpha = 0.4$ is equal to $12.5 \cdot 0.4 = 5$ (see eqn (2)).

The combustion value of the diesel fuel is $Q_{low}^p = 41800 \text{ kJ/(kg of diesel fuel)}$. The boiling temperature is $t_1 = 380^\circ\text{C}$.

According to preliminary calculations for the ECG to achieve the electrical power level of 10 kW the synthesis gas required with the heat of combustion $q_{chem} = 4872 \text{ kJ/(kg of synthesis gas)}$ is

$$B_{sg} = Q_e / (q_{chem} \eta^*), \quad (5)$$

where the coefficient of efficiency (CE) of the ECG is tentatively assumed to be equal to according to the data presented in [26].

For the production of B_{sg} it is necessary to introduce into the burner the following amount of diesel fuel, kg/s

$$G_{df} = B_{sg} \cdot 114/800 \quad (6)$$

and primary air, kg/s

$$G_{air1} = G_{df} \cdot 686.4/114 \quad (7)$$

Heat balance equation for the catalytic burner:

$$Q_{low}^p \cdot \alpha \cdot 114 / 800 (1 - q_3 - q_5) + q_{air1} + q_v = C_c t \quad (8)$$

On the left side of eqn (8) is the heat input from incomplete combustion with taking into account losses with chemical underburning and external heat losses through thermal insulation $Q_{low}^p \cdot \alpha \cdot 114 / 800 (1 - q_3 - q_5)$, from primary air heated in WHB $-q_{air1}$ and steam $-q_v$; on the right side – heat consumption for heating the synthesis gas conversion products, $q_3 = 0.05$, $q_5 = 0.3$.

The coefficient $114/800$, which converts $Q_{low}^p \text{ KJ/(kg of diesel fuel)}$ to $\text{kJ/(kg of synthesis gas)}$, has been found from eqn (2).

The heat introduced by primary air into the burner heated to $t_1 = 380^\circ\text{C}$:

$$q_{air1} = C_{air} t_1 / 800 = 686 \cdot 1.056 \cdot 380 / 800 = 344 \text{ kJ/(kg of synthesis gas)}$$

The heat introduced by vaporized diesel fuel:

$q_v = C_v t_1 / 800 = 3.5380114 / 380 \cdot 189 \text{ kJ/(kg of synthesis gas)}$, where C_{air} , C_v , true specific isobaric heat capacities of air and vaporized diesel fuel at 380°C , are equal to 1.056 and 3.5 kJ/(Kkg) , respectively.

True specific isobaric heat capacity of synthesis gas $C_{sg} = 1.62 \text{ kJ/(K kg of synthesis gas)}$.

From the heat balance equation, the temperature in the catalytic burner can be determined as:

$$t = \left[Q_{low}^p a(1 - q_3 - q_5) 114 / 800 + q_{air1} + q_v \right] / C_{sg} \quad (9)$$

Chemical efficiency of the catalytic burner can be written as:

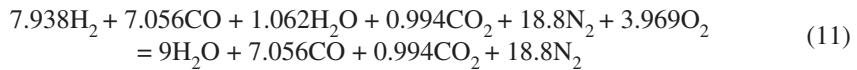
$$\eta_{chem} = B_{sg} \cdot q_{chem} / (G_{dt} \cdot Q_{low}^p) \quad (10)$$

2.2.2 Synthesis gas in the electrochemical generator: physicochemical processes

Synthesis gas cooled in the heat exchanger (4) down to $t_3 = 800^\circ\text{C}$ is fed to the anode channel, while the air heated up to $t_2 = 620^\circ\text{C}$ in the heat exchanger (4) is fed to the cathode channel of the ECG. According to [27], contribution to the EMF value from the CO oxidation reaction at the SOFC anode is less than 1% of that from hydrogen oxidation; therefore, it is neglected in the calculations, and assumed that only hydrogen is oxidized at the SOFC anode.

The water-gas reaction $\text{CO} + \text{H}_2\text{O} \rightleftharpoons \text{CO}_2 + \text{H}_2$ cannot be a donor of hydrogen, since this reaction proceeds only on the iron-chromium catalyst [27], which is absent in the anode channel.

At the SOFC anode, the oxidation of hydrogen from the synthesis gas by oxygen coming from the electrolyte is described by a stoichiometric equation assuming all the hydrogen to be completely oxidized:



The exothermic effect of reaction (11) under standard conditions ΔH_1^0 is equal to 2400 kJ/(kg of synthesis gas), while the change of Gibbs free energy for the reaction under standard conditions $\Delta G^0 = -2863.59$ kJ/(kg of synthesis gas), $\Delta S^0 = 0.44$ kJ/(kg K).

According to [28], the temperature of the hydrogen oxidation products at the anode is

$$T_3^* = (\Delta H_1^0 - \Delta G^0) / \Delta S^0 \quad (12)$$

The temperature level is in accordance with [19]. The EMF of a planar cell at the temperature of T_3^* is equal to

$$E_e = \Delta G / (2F) \quad (13)$$

where $F = 9.648 \cdot 10^4$ C/mol is the Faraday constant; ΔG kJ/(kmol of water) is determined at T_3^* .

Fuel cells usually operate under isobaric-isothermal conditions. The work in the isobaric-isothermal process is equal to the negative change in of the free Gibbs energy [28], and the electrical power is equal to the negative change in the free Gibbs energy ΔG , kJ/(kg of synthesis gas) at T_e^* , multiplied by the synthesis gas consumption B_{sg} , kg/s:

$$Q_e = B_{sg} \Delta G \quad (14)$$

On the other hand, the electrical power is [19]:

$$Q_e = B_{sg} q_{chem} \eta \quad (15)$$

Equating these expressions, the electrical efficiency of the SOFC battery is determined and then compared with [26, 29].

Heat balance equation for the electrochemical unit

It is assumed that the SOFC battery is ideally insulated, the chemical underburning is equal to zero.

The equation for the balance of incoming and outgoing energy flows has the form

$$B_{sg} \phi_T \Delta H_1^0 + B_{sg} C_{sg} t_3 + G_{air} C_{air} t_2 = Q_e + C_{sg}^* B_{sg}^* t_3^* + C_{air}^* G_{air}^* t_2^* \quad (16)$$

where:

$B_{sg} \phi_T \Delta H_1^0$ – power, selected the oxidation of hydrogen at the anode SOFC,

$B_{sg} C_{sg} t_3$ – power paid synthesis-gas emerging from the cooler (see Fig. 1),

$G_{air} C_{air} t_2$ – power, the air coming from the boiler-heat exchanger,

Q_e – electrical power removed from the ECG,

$B_{sg}^* C_{sg}^* t_3^*$ – power assigned to partially oxidized synthesis-gas, partly oxidized, outgoing from anode channel,

$G_{air}^* C_{air}^* t_2^*$ – power, assigned to the air outgoing from the cathode canal,

ϕ_T – a portion of hydrogen oxidized at the anode.

The temperature t_3 at the inlet of the synthesis gas to the anode channel is assumed to be 800°C. The specific true isobaric heat capacity of the synthesis gas at 800°C is determined from eqn (2) as $C_{sg} = 1.5$ kJ/(K kg of synthesis gas).

The air temperature at the outlet from the synthesis gas cooler (at the inlet to the cathode channel) t_2 is determined from the heat balance equation of the synthesis gas cooler and equal to 620°C. The specific true isobaric heat capacity of the air at the temperature of 620°C $C_{air} = 1.115$ kJ/K kg (found from eqn (3)). The temperature t_3^* is equal to 780°C (it was defined earlier, see Form. (12)).

The specific true isobaric heat capacity of the oxidized synthesis gas at the outlet from the anode at 780°C $C_{sg}^* = 1.161$ kJ/(K·kg of synthesis gas) is determined from eqn (11). The air temperature at the outlet from the cathode channel is assumed to be by 100°C below, i.e. $780 - 100 = 680^\circ\text{C}$.

The specific true isobaric heat capacity of the air depleted in oxygen at the outlet from the cathode channel is $(7.497 - 3.969)\text{O}_2 + 28.18\text{N}_2$ at 680°C, $C_{air}^* = 1.166$ kJ/Kkg.

The air flow rate into the cathode channel necessary for the complete oxidation of synthesis gas can be calculated using eqn (3) as:

$$G_{air} = (240 + 789) B_{sg} / 800 = 1.285 \cdot 5.32 \cdot 10^{-3} = 6.84 \cdot 10^{-3} \text{ kg/s.}$$

The flow rate of vitiated air at the outlet of the cathode channel can be determined from eqns (3) and (11):

$$G_{air}^* = G_{air} - 3.969 \cdot 32 \cdot B_{sg} / 800 = (6.84 - 0.158 \cdot 5.32) \cdot 10^{-3} = 5.99 \cdot 10^{-3} \text{ kg/s}$$

The consumption of synthesis gas at the outlet of the anode channel can also be determined from eqn (11):

$$B_{sg}^* = (9 \cdot 18 + 7.656 \cdot 28 + 0.994 \cdot 44 + 18.8 \cdot 28) B_{sg} / 800 = 1.16 \cdot 5.32 \cdot 10^{-3} = 6.17 \cdot 10^{-3} \text{ kg/s}$$

The materials balance at the ECG input and output has converged.

From the heat balance eqn (16), a portion of hydrogen oxidized at the anode can be calculated

(all the parameters included in the heat balance equation are presented in Table 1):

$$\varphi_T = \left(Q_e + C_{sg}^* B_{sg}^* t_3^* + C_{air}^* G_{air}^* t_2^* - B_{sg} C_{sg} t_3 - G_{air} C_{air} t_2 \right) / \left(B_{sg} \Delta H_1^0 \right) \quad (17)$$

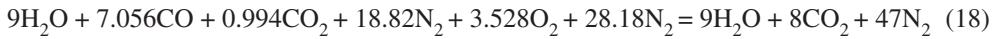
Substitution in the last equation of the parameters presented in Table 1 gives the value of φ_T , which agrees with the data of [30].

2.2.3 Modelling of the processes in the waste-heat boiler – utilizer: physicochemical processes

The rest hydrogen, not oxidized at the anode, is oxidized in WHB with power release:

$$(1 - \varphi_T) B_{sg} \Delta H_1^0$$

Carbon oxide, coming from the anode channel, is oxidized by oxygen-depleted air (7.497–3.969) O₂ + 28.18N₂, outgoing from the cathode channel, according to the stoichiometric equation:



The exothermic effect of reaction (18) is kJ/(kg of synthesis gas).

Power balance for the waste-heat boiler-utilizer.

The equation for the balance for the waste-heat boiler-utilizer has the form

$$\begin{aligned} & (1 - q_3 - q_5) \left(C_{sg}^* B_{sg}^* t_3^* + C_{air}^* G_{air}^* t_2^* + (1 - \varphi_T) B_{sg} \Delta H_1^0 + \Delta H_2^0 B_{sg} \right) \\ & = G_{air1} C_{air} (t_1 - t_{air}) + G_{df} \left[C_{df} (t_1 - t_{df}) + r_{df} \right] + G_{ex} C_{ex} t_{ex} \end{aligned} \quad (19)$$

Table 1: Values of the parameters in eqn (17).

Value	Unit	Designation
5.32·10 ⁻³	kg/s	B _{sg}
6.17·10 ⁻³	kg/s	B _{sg} [*]
1.5	kJ/K·kg	C _{sg}
1.161	kJ/K·kg	C _{sg} [*]
1.166	kJ/K·kg	C _{air} [*]
1.115	kJ/K·kg	C _{air}
6.84·10 ⁻³	kg/s	G _{air}
5.99·10 ⁻³	kg/s	G _{air} [*]
620	°C	t ₂
680	°C	t ₂ [*]
800	°C	t ₃
780	°C	t ₃ [*]
2400	kJ/K·kg	ΔH ₁ ⁰
10	kW	Q _e

Where:

- $C_{sg}^* B_{sg}^* t_3^*$ – the power coming with products from the anode channel,
- $C_{air}^* B_{air}^* t_2^*$ the power coming with air from the cathode channel,
- $(1-\phi_T) \cdot B_{sg} \Delta H_1^0$ – the power coming released from the oxidizing the rest hydrogen,
- $\Delta H_2^0 B_{sg}$ – the power non-oxidized at the anode and carbon oxide, with primary air,
- $G_{air1} C_{air} t_{air}$ – the power inflowing to the catalytic burner,
- $G_{df} C_{df} t_{df}$ – the power with diesel fuel,
- $G_{air1} C_{air} t_1$ – the power is spent on heating the primary air,
- $G_{df1} C_{df} (t_1 - t_{df}) + r_{df}$ – heating to boiling and evaporating the diesel fuel,
- $G_{ex} C_{ex} t_{ex}$ – heating exhausted combustion products from the boiler-utilizer, $q_3 = 0.03$, $q_5 = 0.242$.

From the heat balance equation, the temperature of the gases outgoing from the WHB can be determined (all the parameters in eqn (19) are presented in Table 2).

The specific isobaric heat capacity of diesel fuel is $C_{df} = 1.84 \text{ kJ/K} \cdot \text{kg}$.

$$G_{ex} = B_{sg}^* + G_{air}^* = B_{sg} + G_{air} = (6.17 + 5.99)10^{-3} = (5.32 + 6.84)10^{-3} = 12.16 \cdot 10^{-3} \text{ kg/s}$$

Table 2: Values of the parameters in eqn (20).

Value	Unit	Designation
$5.32 \cdot 10^{-3}$	kg/s	B_{sg}
$6.17 \cdot 10^{-3}$	kg/s	B_{sg}^*
1.161	kJ/K·kg	C_{sg}^*
1.166	kJ/K·kg	C_{air}^*
1.056	kJ/K·kg	C_{air}
1.38	kJ/K·kg	C_{ex}
$4.56 \cdot 10^{-3}$	kg/s	G_{air1}
$5.99 \cdot 10^{-3}$	kg/s	G_{air}^*
1.84	kJ/K·kg	C_{df}
$0.758 \cdot 10^{-3}$	kg/s	G_{df}
$12.16 \cdot 10^{-3}$	kg/s	G_{ex}
380	°C	t_1
20	°C	t_{air}
680	°C	t_2^*
780	°C	t_3^*
20	°C	t_{df}
2400	kJ/kg	ΔH_1^0
2472	kJ/kg	ΔH_2^0
0.745	Rel. unit	ϕ_T
644	kJ/kg	r_{df}

r_{df} – the heat of evaporation of diesel fuel kJ.kg.

$$t_{ex} = \left\{ (1 - q_3 - q_5) \left(C_{sg}^* B_{sg}^* t_3^* + C_{air}^* G_{air}^* t_2^* + (1 - \varphi_T) B_{sg} \Delta H_1^0 + \Delta H_2^0 B_{sg} \right) - G_{air1} C_{air} (t_1 - t_{air}) - G_{df} \left[C_{df} (t_1 - t_{df}) + r_{df} \right] \right\} / (G_{ex} C_{ex}) \quad (20)$$

The results of the calculations are summarized in Table 3.

Power supplied to the network heater to heat the network water for hot water needs of heat consumers can be calculated as

$$Q_{nh} = G_{ex} C_{ex} (t_{ex} - t_{ex}^*) \quad (21)$$

where t_{ex}^* is the temperature after the network heater, which is assumed to be equal to 120°C.

Coefficient of fuel utilization in the mini-CHPP can be calculated as follows:

$$\eta_{ut} = (Q_e + Q_{nh}) / (Q_{low}^p G_{df}) \quad (22)$$

Table 3: Results of calculations for the mini-CHPP typical operation.

Parameter designation	Designation	Dimension	Eq. number	Value
Air flow rate at the sooting formation point	α_c	dim. less	(1)	0.32
Synthesis gas flow rate (specific consumption)	B_{sg}	kg/s	(5)	$5.32 \cdot 10^{-3}$
Diesel fuel flow rate (specific consumption)	G_{df}	kg/s	(6)	$0.758 \cdot 10^{-3}$
Primary air flow rate to the burner	G_{air1}	kg/s	(7)	$4.56 \cdot 10^{-3}$
Burner temperature	t	°C	(9)	1284
Burner chemical efficiency	η_{chem}	%	(10)	81.7
Temperature at the SOFC anode	T_3^* t_3^*	K °C	(12)	1053 (780)
EMF of the single planar cell	E_e	V	(13)	0.985
ECG electrical power	Q_e	kW	(14)	10
ECG electrical efficiency (gross)	η	(%)	(15)	0.386
Volume fraction of hydrogen oxidized at the anode	φ_T	%	(17)	74.5
Temperature of gases outgoing from the boiler-utilizer	t_{ex}	°C	(20)	950
Network water heater capacity	Q_{nh}	kW	(21)	13.9
Coefficient of fuel utilization in CHPP	η_{ut}	dim. less	(22)	0.754

Absolute consumption of diesel fuel for the production of electrical power	G_e	kg/h	(23)	1.05
Specific consumption of diesel fuel for the production of electrical power	b_e	kg/kW·h	(24)	0.114
Absolute consumption of diesel fuel for the production of heat power	G_T	kg/h	(25)	1.587
Specific consumption of diesel fuel for the production of heat power	b_T	kg/GJ	(26)	31.7
Difference in diesel fuel flow rates supplied to ECG and used beneficially in power plant	ΔG_{df}	kg/s	(27)	0
Reagents input flow rate	G_{in}	kg/s	(28)	$12.16 \cdot 10^{-3}$
Reagents output flow rate	G_{out}	kg/s	(29)	$12.16 \cdot 10^{-3}$
Flow rate imbalance	ΔG	kg/s	(30)	0
Power input	Q_{in}	kg/s	(31)	25.94
Power output	Q_{out}	kg/s	(32)	25.94
Power imbalance	ΔQ	kg/s	(33)	0

3 ENERGY PERFORMANCE OF THE MINI-CHPP

3.1 Differentiation of the diesel fuel consumption between the production of electrical and heat energy and the specific fuel consumption

Below are calculated in line with the methodology [31], absolute and specific consumption of diesel fuel in cogeneration of electricity and heat production.

Absolute consumption of diesel fuel for the production of electrical power, kg/h

$$G_e = Q_e \cdot 1 \cdot 3600 / (Q_{low}^p \eta_{ut}) \quad (23)$$

Specific consumption of diesel fuel for the production of electrical power, kg/kW·h

$$b_e = G_e / (Q_e \cdot 1). \quad (24)$$

Absolute consumption of diesel fuel for the production of heat power, kg/h

$$G_T = Q_{nh} \cdot 1 \cdot 3600 / (Q_{low}^p \eta_{ut}) \quad (25)$$

Specific consumption of diesel fuel for the production of heat power, kg/kW·h

$$b_T = G_T \cdot 10^6 / (Q_{nh} \cdot 1 \cdot 3600) \quad (26)$$

Imbalance of the diesel fuel utilization, kg/h

$$\Delta G_{df} = G_{df} - (G_e + G_T) / 3600 \quad (27)$$

The levels of diesel fuel specific consumption correspond to a high-efficient heat power-station [32].

3.2 Material balance of the electric power plant.

Material balance is determined by the equity consumptions of fuel, oxidizer and the products of combustion at the input and output of the mini CHP.

$$\text{Input flow rate } G_{in} = G_{air} + G_{air1} + G_{df} \quad (28)$$

$$\text{Output flow rate } G_{ex} = B_{sg}^* + G_{air}^* \quad (29)$$

$$\text{Mass flow rate imbalance } \Delta G = G_{in} - G_{out} \quad (30)$$

3.3 Heat balance of the electric power plant.

Heat balance is determined by the equity capacity energy flows at the input and output, taking into account the heat loss of the installation.

$$\text{Input power } Q_{in} = Q_{low}^p G_{df} \eta_{chem} \quad (31)$$

$$\text{Output power } Q_{out} = Q_e + G_{ex} C_{ex} t_{ex} \quad (32)$$

$$\text{Power imbalance } \Delta Q = Q_{in} - Q_{out} \quad (33)$$

4 CONCLUSIONS

The principal possibility for using the products of the motor diesel fuel air conversion for the operation of the SOFC-based electrochemical generator has been demonstrated.

Calculation studies have shown that the heat of combustion of the products of the diesel fuel air conversion (at $\alpha = 0.4$) is equal to 4872 kJ/(kg of synthesis gas), the fuel cell EMF is 0.985 V, the portion of hydrogen oxidized at the anode is 0.894, the temperature of hydrogen oxidation at the SOFC anode is 780 °C. The results of calculations carried out for the ECG are in reasonable agreement with the literature data obtained for other liquid hydrocarbon fuels.

Comparison with modern power sources has shown that both the specific consumption of diesel fuel for the production of electric power equal to 114 g/kWh (162 g r.f./kW·h) and heat power equal to 31.7 kg/GJ (45.1 kg r.f./GJ, 189 kg r.f./ Gcal) corresponds to a high-efficient heat power-station and lower than in modern diesel-electric stations of equal power.

It is noted that thermal energy for heating and hot water supply in this cogeneration technology is an additional product that reduces overall consumer costs.

LIST OF ACRONYMS

- CE – Coefficient of Efficiency
- SG – Synthesis Gas
- WHB – Waste-Heat Boiler-utilizer
- SOFC – Solid Oxide Fuel Cell
- CHPP – Combined Heat and Power Plant

- EMF – Electro Motive Force
- ECG – ElectroChemical Generator
- GTP – Gas-Turbine Plant
- STP – Steam-Turbine Plant

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